

## Standardization of Sodium Thiosulphate using iodometric titration



### General Aim

Standardization of a solution of sodium thiosulphate

### Method

Iodometric titration

### Learning Objectives (ILOs)

- Standardization of a prepared solution of sodium thiosulphate.
- Understanding the difference between iodometry and iodimetry.
- Understanding the concept of iodometry.
- Understanding the concept of redox reactions.

### Theoretical Background/Context

- **Iodometric titration:** The titration in which the sample has oxidizing properties. Upon adding potassium iodide, free iodine is liberated equivalent to the sample. The liberated iodine can be easily determined against a standard reducing agent.
- **Iodimetric titration:** The titration in which the sample has reducing properties and thus can be determined directly against iodine.

### Principle of Work

Sodium thiosulphate is a reducing agent which can be standardized against iodine (oxidizing agent) by iodometric titration.